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Acids, Bases and

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~~Activity 2.2 |
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Uses of Acids

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Neutralization

Reaction Of

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Unit 4 (page
109?182) We will
cover the
following

topics: 1.

Amphiprotic

Substance 2.

Amphoteric

Compounds CHEM

0012 Lecture

Notes 17 Acids,

Bases and Salts

Hebden - Unit 4

(page 109?182)

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1. Amphiprotic
Substance A
substance that
can act either
as a proton
acceptor or a
proton donor.

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pronouncement
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while he cooks
up something
special. In this
interactive
tutorial you
will compare and
classify the
properties of
compounds that
are acids,
bases, and salts

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and identify
basic examples
of these
compounds.

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on Thursday May
10, 2012. Terms
in this set (43)
Acids taste

sour, will
change the color
of an acid-base
indicator, and
can be. strong
or weak
electrolytes in
aqueous
solution.

Chapter 19

Page 22/52

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Chapter 19 Acids
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Let's learn more
about Acids

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Bases and Salts!

We will discuss
the important
properties of

Acids, Bases and
Salts and how
they differ from
each other.

Simple examples
and interesting
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answers & doubts
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Kiddy Math

A salt is an ionic compound that can be

formed by the neutralization reaction of an acid and a base.

The following table shows the patterns for the solubility of salts . To help you remember,

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think of SPANE

for soluble
salts: Sodium,
Potassium,

Ammonium,

Nitrates and

Ethanoates.

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A substance used
as an acid-base
indicator. (6)

11. An atom or

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molecule with a positive or negative charge.

(3) 13. The strength of a base. (10) 15.

Common name for sodium

hydroxide. (3)

17. Color of litmus paper in acids. (3)

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Compounds formed
from

combinations of
acids and bases
taste salty and
are referred to
in chemistry as
salts. Sweet
compounds have

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Bases And
characteristics
of both acids
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and bases in the
same molecule.

We will explore
the relationship
between
molecular
structure and
acids-bases, and
consider water
solutions of
acids and bases.

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Chemistry ...

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Notes MODULE - 2

Matter in our
Surroundings

Acid Base INTEXT
QUESTION 8.1 1.

Put the

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substances in
acid or base
bottle. (a) Milk

of magnesia (b)
gastric juice in
humans (c) soft
drinks (d) lime
water (e)
vinegar (f) soap

2. What will
happen if you
add a drop of
the following on

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Notes ACIDS,
BASES AND SALTS
Bronsted-Lowry
base = proton
acceptor (H^+ =
proton) Bronsted-
Lowry acid-base
reaction =
reaction
involving the
transfer of a

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proton e.g. HCl

+ NaOH → NaCl +

H₂O H⁺ donates

H⁺ → acid OH⁻

accepts H⁺ → base

e.g. HNO₃ + NH₃

→ NH₄⁺ + NO₃⁻

H⁺ donates H⁺

→ acid NH₃

accepts H⁺ →

base TASK 1 -

Bronsted-Lowry

acids & bases

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uk. A. Answer

the following

questions in

short: 1. State

differences

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Bases And acids

and bases. 2.

Ammonia is found
in many

household

products, such

as window

cleaners. It

turns red litmus

blue. What is

its nature?

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includes a

neutralization

reactions

worksheet and

the

corresponding

key. The

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requires each

student to

balance,

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identify the
formula and
phase of the
salt formed, and
name the acid,
base, and salt
in 18
neutralization
reactions. The
worksheet may
take up to an
entire class per

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