

Chapter 6 Chemical Bonding Section 2 Covalent Answer Key

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Modern Theories of Covalent Bonding
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Chemical Bonding Section

CHAPTER 6 REVIEW Chemical Bonding
SECTION 1 SHORT ANSWER Answer

the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

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Chapter 6: Chemical Bonding Section 1- Introduction to Chemical Bonding

Objectives: define chemical bond;
differentiate between covalent and ionic
bonding; explain why bonding occurs; use
the...

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to Chemical Bonding section 2 Covalent
Bonding and Molecular Compounds
section 3 Ionic Bonding and Ionic
Compounds section 4 Metallic Bonding
section 5 Molecular Geometry CHAPTER
6. Chemistry HMDScience.com Premium
Content Introduction to Chemical Bonding
Key Terms chemical bond nonpolar-
covalent bond ...

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Chemical Bonding Section

A chemical bond is a mutual electrical attraction between the nuclei and valence electrons of different atoms that binds the atoms together. When atoms form a chemical bond, their valence electrons are redistributed to make the atoms more stable. The way the electrons are redistributed determines the type of bond. Chemical bonding that results from the electrical

CHAPTER 6 Chemical Bonding

Modern Chemistry 1 Chemical Bonding

CHAPTER 6 Chemical Bonding

SECTION 1 Introduction to Chemical

Bonding OBJECTIVES 1. Define

Chemical bond. 2. Explain why most atoms form chemical bonds. 3. Describe ionic and covalent bonding.. 4.

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CHAPTER 6 REVIEW Chemical Bonding

SECTION 3 SHORT ANSWER Answer

the following questions in the space provided.

1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom.

2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) positive ions. (d) negative ions.

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Chapter 6 – Chemical Bonds. Jennie L. Borders. Standards. SPS1. Students will investigate our current understanding of the atom. b. Compare and contrast ionic and covalent bonds in terms of electron movement. SPS2. Students will explore the nature of matter, its classification and its system for naming types of matter.

Chapter 6 – Chemical Bonds

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Answer the following questions in the space provided

1 a The notation for sodium chloride, NaCl , stands for one (a) formula unit (c) crystal (b) molecule (d) atom

2 d In a crystal of an ionic compound, each cation is ...

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REVIEW CHEMICAL BONDING
SECTION 3 SHORT ANSWER
ANSWER THE FOLLOWING
QUESTIONS IN THE SPACE
PROVIDED 1 A THE NOTATION FOR
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